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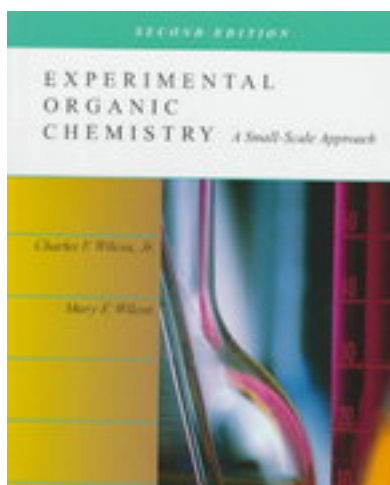
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## Organic Chemistry II Laboratory

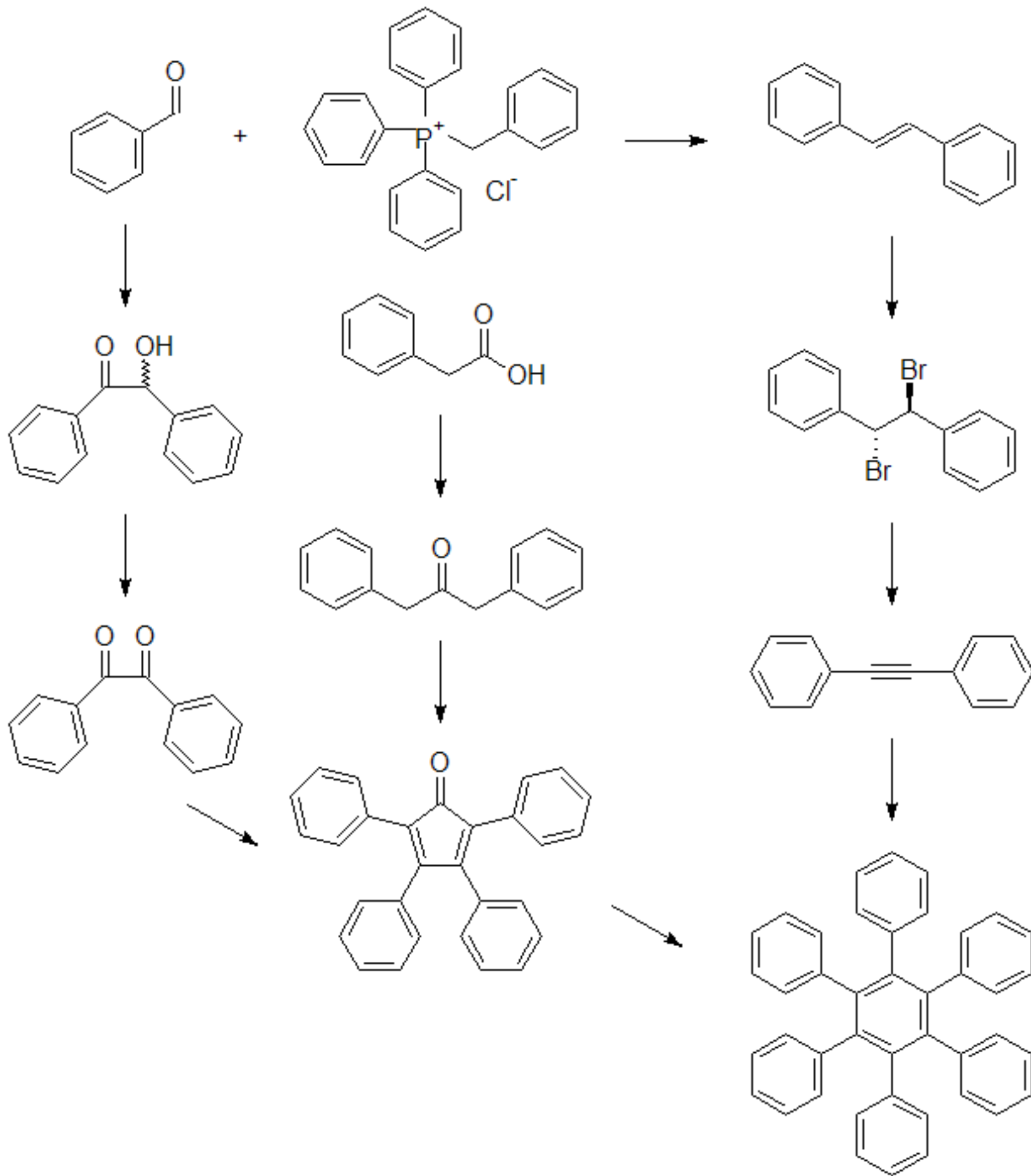
Fall 2018

Text: Experimental Organic Chemistry A Small-Scale Approach, 2<sup>nd</sup> Edition, Wilcox & Wilcox.



**Please note that the scale of each experiment may be changed.**

Instructors: Michael, Sharma, Zhu



Week	Date (Note that our week begins on Tuesdays except for the last week!)	Activity
1	August 28-September 3	No TAs-No Labs
2	September 4-10	Laboratory Safety
3	September 11-17	Benzaldehyde page 75
4	September 18-24	Stilbene page 369
5	September 25-October 1	Stilbene continued

6	October 2-8	Stilbene Dibromide page 370
7	October 9-15	Diphenylacetylene page 370
8	October 16-22	Benzoin page 481-IR Review
9	October 23-29	Benzoin continued-NMR Review
10	October 30-November 5	Benzil page 482
11	November 6-12	Dibenzyl Ketone- <a href="#">Davis, R; Schultz, H.P. Journal Organic Chemistry 1962, volume 27, page 854</a>
12	November 13-19	Tetraphenylcyclopentadienone page 400
13	November 20-25	Holiday-No Labs
14	November 26-December 1	Hexaphenylbenzene page 438
15	December 3-8	Holiday-No Labs

Required materials (You will be turned away and earn a zero in attendance for missing safety attire!):

1. The book
2. Goggles (**not safety glasses**) A seal should form around your eyes.
3. Pants (not shorts) Your legs must be covered.
4. Shoes (not sandals) Your feet must be completely covered.
5. A lab coat

What you should be learning:

1. Know the dangers of each laboratory including how to mitigate your risk.
2. Read and understand all the chapter and not just the experimental procedure.
3. Make sure that you can draw all structures and mechanisms for each lab.
4. Be able to apply a given mechanism to other reagents because we are not just cooks!
5. Understand which fundamental mechanism you are applying (Addition, Elimination or Substitution) including the stereochemical consequences.
6. Understand whether a reaction is an oxidation, reduction or not a net redox reaction.
7. Be able to calculate the yield of a reaction.
  1. structure to formula conversion,
  2. formula to molecular weight conversion,
  3. ml to grams via density and vice-versa,
  4. grams to moles,
  5. what is the limiting reagents,
  6. what is the stoichiometry of the reaction.
8. Understand why a particular procedure was followed.

Your grade will consist of:

1. 1/3 attendance. You must participate fully in the lab, and not come and go as you please!
2. 1/3 laboratory Quizzes. Quizzes will open the day before your lab day (one day).
3. 1/3 laboratory Reports. Reports will open the day after your lab day and will close the day before your

next lab day (six days).

A > 89.5 %, B > 79.5 %, C > 69.5 %, D > 59.5 %

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